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INVESTIGATION OF ELECTRICAL PROPERTIES HOMOGENEITY OF Li-Ti-Zn FERRITE CERAMICS

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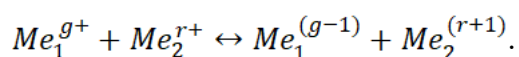
In this article the distribution of electrical conductivity homogeneity of lithium ferrite ceramics was studied. Investigations of the conductivity were performed on samples of lithium substituted ferrite with addition of zirconium dioxide (0%; 0.2%; 0.3%; 0.5 weight %). Experimental samples were sintered at 1010 °C for 2 hours by using the standard ceramic technology. Temperature dependences of electrical conductivity were obtained by the two-probe method (spreading resistance analysis). According to spreading resistance analysis, it was found that the distribution of electrical conductivity in the surface layers varies depending on the chemical composition and this distribution has inhomogeneous character. Also, the activation energy and measurement error were calculated.

Keywords: lithium ferrites, ceramics, electrical conductivity, zirconium dioxide.

Introduction

Ferrite ceramics have a wide range of applications and are used in the most modern electronic computing and radio devices [1, 2]. One of those materials are lithium ferrites, which are widely used in technique that is used both at radio frequencies and at ultrahigh frequencies (microwave range), as well as like cathodes of lithium batteries [3, 4]. The high demand for these materials is due to the low cost of this class of ferrites and excellent ferromagnetic properties. For example, high values of the Curie temperature and saturation magnetization [5], beside low values of dielectric losses [6]. These properties are important for creating ferrite devices with magnetic memory [7].

Spinel of lithium ferrite contains reducible iron ions and lithium ions therefore it is expected that this material has both electronic and ionic conductivity [8]. Experimental data show the electrical conductivity of ferrites increases significantly when ions with different valences are located in equivalent crystallographic positions. The electrical conductivity of ferrites is explained by the Verwey mechanism [9], which means that electrons hop from ion to ion. In this case, the valence of the ions changes in accordance with the expression:



In case when ferrite has the same ions in equivalent crystallographic lattice sites, which has differ by one, then both ions exchange their valences. The corresponding valence states migrate along the crystal, and if it has a sufficiently high concentration of ions with different valences so it leads to high conductivity.

From the Verwey mechanism of electrical conductivity discussed above, it follows that the conductivity of ferrites is more dependent on the content of multivalent ions in them. Such ions in ferrites are often iron ions, as well as ions of zinc, manganese, titanium, and some others [10].

The ceramic method is a classic method of manufacturing lithium ferros spinels, which includes a synthesis stage with the formation of a single-phase product and a further sintering stage at high temperatures [11-13]. However, during high-temperature annealing, the capabilities of this method are limited by the low thermal stability of some reagents [14] (evaporation of lithium and zinc oxide) of the initial mixture and their incomplete ferritization. For this reason, it is observed

probability of increasing of adverse phase inclusions and decreasing of the chemical and product structural homogeneity. These facts lead to a deterioration in its structural, electrical and magnetic properties, which increases the yield of defective products [15, 16].

Conventional methods for solving these problems are using of additional technological methods, including operations of multiple grinding, briquetting and subsequent firing of reaction mixtures. At the same time, such approaches are extremely laborious, energy consuming, and difficult to implement due to the multiple increase in intermediate operations. In addition, the appearance probability of impurities (contaminants) in the reaction mixture is increased.

Effective methods to reduce sintering temperature include such methods as adding oxides with a low melting point, ball mill grinding [17–20], sol-gel [21], and other technologies [22–24]. It was shown in [25, 26] that heating reagents with intense electron beams is an effective method of intensifying solid-state reactions [27], leading to a decrease in the synthesis temperature and an increase in the homogeneity of ferrites [28].

Despite the availability of ferrite production around the world, there are still problems associated with the need to improve the quality of products. In the same time, the development and improvement of electronic technology is accompanied by an ever-increasing tightening of requirements for magnetic materials. Today, there is an urgent necessity for highly specialized magnetic materials with a certain combination of electro-physical and magnetic properties. Alloying lithium and lithium substituted ferrites with different additives allow achieving the necessary characteristics or improving existing electrical, magnetic and mechanic characteristics. It allows get almost any properties inherent in more expensive materials.

Several types of added additives can be distinguished: ones of those form a liquid phase during sintering, thereby affecting the ferrite microstructure, others are located at the grain boundaries as the second phase, which has a very high resistivity. In addition, some additives can be integrated into the ferrite lattice and replace ions at tetrahedral or octahedral levels.

Recently, zirconium dioxide has become an effective additive allowing one to influence the properties and structural state of a wide class of materials [29-31]. Previous scientific results show the additive affects the structure of ferrites, which makes it possible to control other structurally sensitive properties, such as electrical conductivity [32], magnetic permeability [33], thermal conductivity [34], and the shape of the hysteresis loop.

Therefore, there is every reason to suppose that the introduction of zirconia additives can be used in the manufacture of lithium ferrites. This can effect on the formation of structural and electromagnetic characteristics and will allow obtain improved properties compared with prototypes. Thus, the homogeneity of the distribution of electrical conductivity in the surface layers of multi-component lithium ferrite (with the chemical composition $\text{Li}_{0.65}\text{Fe}_{1.6}\text{Ti}_{0.5}\text{Zn}_{0.2}\text{Mn}_{0.05}\text{O}_4$) was studied. The effect of the introduction of zirconia additives on the electrical properties of this ferrite is considered.

1. Experimental part

The measurements were carried out on samples of Li-Ti-Zn ferrite with a low contain of ZrO_2 additive. The initial reagents were chemically pure powders of zinc oxide (ZnO), manganese oxide (MnO), zirconium dioxide (ZrO_2), titanium dioxide (TiO_2) and iron oxide (Fe_2O_3), as well as lithium carbonate (Li_2CO_3) [32, 33]. All initial reagents were pre-dried in a laboratory furnace at a temperature of 200°C for 240 minutes.

In accordance with the chemical equation ($\text{Li}_{0.65}\text{Fe}_{1.6}\text{Ti}_{0.5}\text{Zn}_{0.2}\text{Mn}_{0.05}\text{O}_4$), the initial reagents were weighed on Shimadzu AUW-D digital analytical balances. Then powders were mixed in an agate mortar with tenfold rubbing through a metal sieve with a mesh size of $100\ \mu\text{m}$. It was done to achieve the most uniform distribution of the initial reagents. The mixture of initial powders was divided into four equal parts, after that the zirconia (ZrO_2) was added in different weight ratios to each part. The weight content of ZrO_2 additive was 0; 0.2; 0.3; 0.5 wt. %.

Each part of the mixture was mechanically activated in EMAX (Retsch) high-speed ball mill to increase the homogeneity of the distribution of zirconium dioxide in the mixture and its reactivity [34-35]. Mechanical activation was carried out in steel grinding vials and balls with a diameter of 2 mm. The weights of the mixture and balls were in the ratio 1:10. The mechanical activation was performed at 2000 rpm for 30 minutes.

Mechanically activated powders were pressed by single-sided cold pressing in a hydraulic hand press PGr-10. The pressing pressure was 200 MPa. The pressing time was 3 minutes. After pressing, the samples had a cylindrical shape with a diameter of 15 mm.

The studied samples were sintered at 1010° C for 120 minutes in a laboratory resistance furnace in an air atmosphere. The X-ray diffraction analysis was carried out by using ARL X'TRA (Switzerland) diffractometer with a semiconductor Si (Li) Peltier detector and Cu K_α radiation. XRD patterns were measured in the range $2\theta = (10-70)^\circ$ with scanning rate of $0.02^\circ \cdot \text{sec}^{-1}$ and were processed by the full profile analysis using the Powder Cell 2.5 software, where the pseudo-Voigt profile function was used. Phases were identified by the PDF-4+ powder database of the International Center for Diffraction Data (ICDD). The electrical characteristics of the samples were measured by the two-probe method [36].

It is possible to record the current change during the heating of sample with using this method. The studying sample is located on a substrate, which is gradually heated to a temperature of 540 K with using a spiral heater. Two probes are fixed on the surface of the sample, one of which is supplied with a voltage of 5V, and the measured current is detected from the other probe with a voltmeter. This method is characterized by a high locality of measurements, since the spreading resistance is determined mainly by the contact area. The volume of the region where the resistivity is determined is approximately 10^{-10} cm^3 . According to the data, which obtained from measurements by the two-probe method, it is possible to calculate the volume resistivity by the equation:

$$\rho = \frac{U \cdot r_0 \cdot \pi}{I}, \quad (1)$$

where U is the applied voltage, I is the current, r_0 is the diameter of the contact.

The diameter of the contact is approximately $\sim 10 \mu\text{m}$ and corresponds to the spatial resolution of the method. For Li-Ti-Zn ferrite samples, temperature dependences of the current on temperature were obtained. Each sample was heated from room temperature to 540°K. To determine the accuracy of the study, the surface current of each sample was measured at four points. According to the temperature dependences of the current, experimental curves were calculated in the coordinates $\ln(I)=f(1/(k \cdot T))$ to determine the activation energy. The activation energy is defined as the tangent of the angle of inclination to the approximation curve [37].

2. Results and discussions

Figure 1 shows X-ray diffraction patterns of Li-Zn-Ti ferrite with a 2% content of ZrO₂ and the initial powder of ZrO₂. Diffraction patterns show all the high-intensity reflections which belongs to zirconium dioxide, coincide with the Li-Zn-Ti ferrite reflections. Only one reflection located at $2\theta \approx 51^\circ$ does not coincide. By the presence of reflections at these angles, it is possible to detect zirconium dioxide in the composition of ferrite, however, the small weight content of ZrO₂ in the mixture, as well as the presence of overlapping reflections, does not allow to detecting its content in the studied samples.

Figure 2 shows the diffraction patterns of the initial powder components Fe₂O₃ (JCPDS No. 40-142), Li₂CO₃ (JCPDS No. 66-941) and TiO₂ (JCPDS No. 82-656) phases.

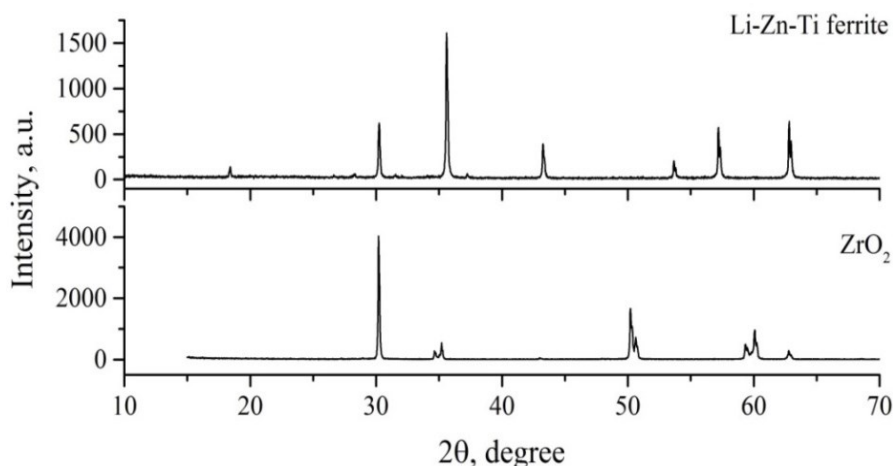


Fig.1. X-ray diffraction patterns of Li-Zn-Ti ferrite and ZrO_2 powder

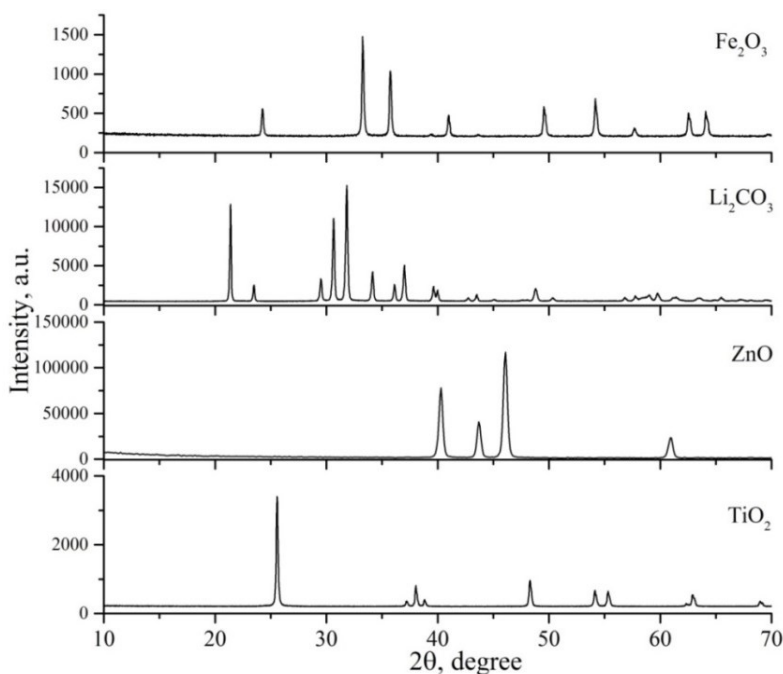


Fig.2. X-ray diffraction pattern of initial powders

Analysis X-ray diffraction patterns shows absences of reflections which can be belong to initial powder and all observed reflection belongs to spinel phase. It shows that the selected technological modes of sintering, pressing, as well as the time and energy of mechanical activation allow us to obtain the spinel phase, which corresponds to final composition of Li-Zn-Ti ferrite without impurities. For all Li-Ti-Zn ferrite samples, current measurements were carried out during the heating of the sample. Values of the current were taken on the surface of the sample at four randomly selected points.

Figure 3 shows the temperature dependences of the bulk conductivity current for a sample without zirconia, which measured at four points on the surface. The temperature curves of the bulk conductivity current are characterized by a linear dependence in logarithmic coordinates $\ln(I)=f(1/kT)$. With increasing temperature, an increase in electrical conductivity is observed, which is characteristic of lithium ferrites.

Figures 4, 5 and 6 show curves for samples with additive of zirconium dioxide (0.2, 0.3 and 0.5wt %). In these graphs, the curves are located as close as possible to each other, which indicate a small experimental error and may indicate a uniform distribution of electrical conductivity over the surface of the sample.

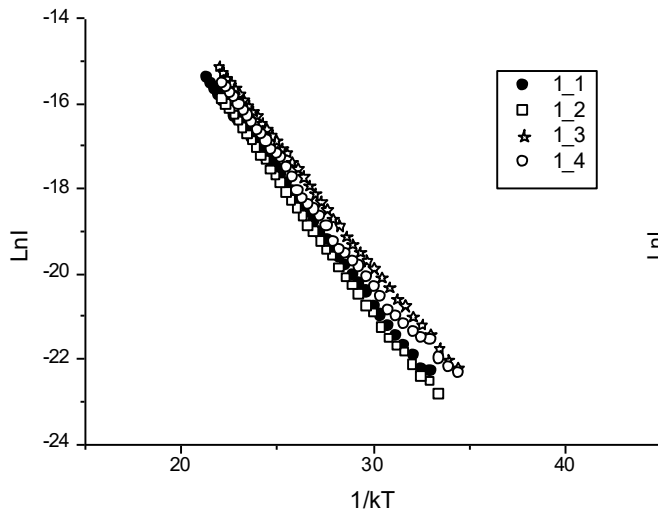


Fig.3. The temperature dependence of conductivity current for samples without ZrO_2

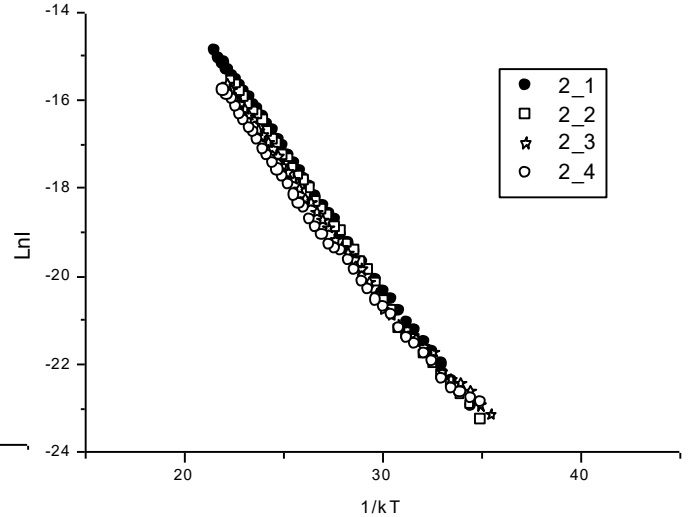


Fig.4. The temperature dependence of conductivity current for Li-Ti-Zn ferrite with 0.2 wt% ZrO_2

The linear form of dependence without kinks in Figures 3–6 indicates that in the temperature range 300–540 °K there is only one conduction mechanism with a certain activation energy. This type of temperature dependence is due to a change in electron mobility in Li-Ti-Zn ferrite, during the heating of the sample. In this case, the Verwey mechanism of conductivity is best suited to describe the mechanism of electrical conductivity.

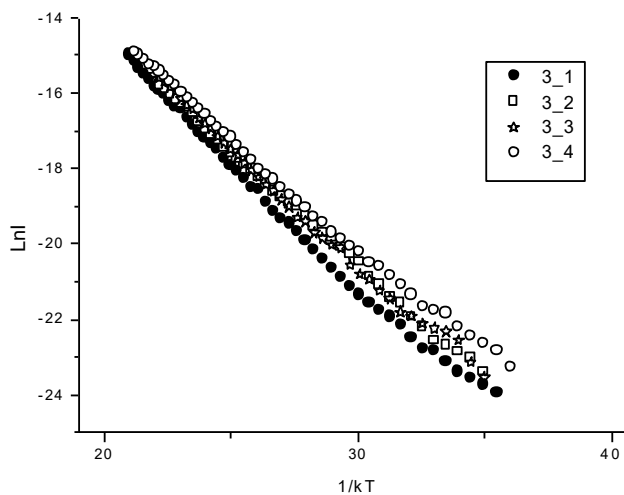


Fig.5. The temperature dependence of conductivity current for Li-Ti-Zn ferrite with 0.3 wt% ZrO_2

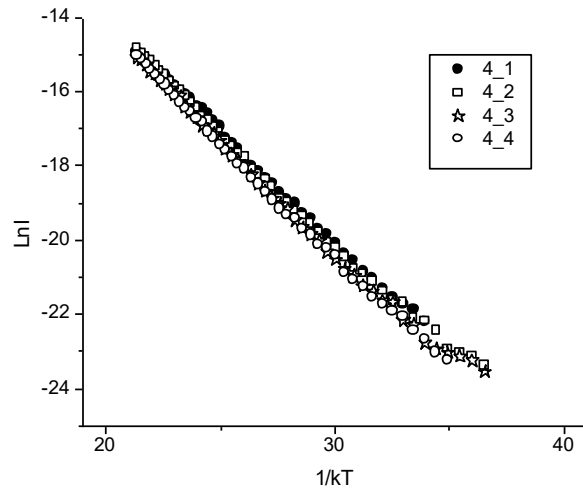


Fig.6. The temperature dependence of conductivity current for Li-Ti-Zn ferrite with 0.5 wt% ZrO_2

This mechanism explains the electrical conductivity of ferrites by the process of electron hopping from ion to ion, so the electrical conductivity will depend on the content of multivalent ions in the ferrite. For Li-Ti-Zn ferrite samples, the electrical conductivity will be determined by electron hopping between the multivalent iron ions Fe^{2+} and Fe^{3+} .

In Figures 3-6, the electrical conductivity of ferrite samples increases during the heating process. The similar character of the dependences $\ln(I) = f(1/kT)$ for samples with and without zirconium dioxide indicates that additive of zirconium dioxide does not significantly effect on the mechanism of electric conductivity in lithium ferrites. A slight change of the electrical characteristics may be due to the effect of zirconium dioxide on the microstructural properties of Li-Ti-Zn ferrite. As it is known, microstructural properties can effect on the electrical characteristics. So, for example, a decrease of grain size and an increase of porosity can affect on the electrical resistance of ferrites.

The experimental temperature curves were used to calculate the activation energy (E_a). The activation energy can be calculated if obtained experimental points fit in a straight line. In this work, the activation energy was calculated for all samples; the average value of the activation energy based on four measurements is presented in Table 1 (column 2). Also, for each measurement, the volume resistivity was calculated according to equation 1, the average value of the resistivity can be seen in table 1 (column 4). From the data of Table 1 it can be seen that the samples have a high volume resistivity. The activation energy and resistivity data, which obtained from the surface of samples with different ZrO_2 additive contents, have an insignificant difference, which indicates that small additions of zirconium dioxide up to 0.5% do not significantly affect the electric conductivity in Li-Ti-Zn ferrite.

Table 1. Electrophysical parameters of Li-Ti-Zn ferrite

Additive of ZrO_2 , %	E_a , eV	ρ , GO m·cm
0	0.597±0.03059	1.45±0.175
0.2	0.594±0.02803	1.34±0.109
0.3	0.604±0.03086	1.58±0.206
0.5	0.593±0.02775	1.27±0.115

According to Table 1, we can conclude that the electrical resistivity is evenly distributed over the surface of the sample; the deviation from the average value of the activation energy for each sample based on four measurements is not more than 7%. So the largest variation of the measurement is the sample without the addition of zirconium dioxide. For this sample, the actual error of the activation energy value exists in the probability interval with a standard deviation of ± 0.03059 eV at an arbitrary temperature of the sample in the temperature range 300-540 °K.

Conclusion

In the presented work, samples of Li-Ti-Zn ferrite with the addition of zirconium dioxide were made by the standard ceramic technology (Conventional method). X-ray phase analysis showed the presence of a spinel phase which corresponds to phase of final Li-Ti-Zn ferrite. This fact indicates a complete ferritization process.

In this article, small additives of zirconium dioxide up to 0.5 wt% were considered, it was found that such a low content of the additive does not has significantly affect on the electrically transfer process in Li-Ti-Zn ferrite. The samples under study possess semiconductor properties; an increase in the conductivity is observed with an increase in the temperature of the samples.

For the studied lithium ferrite, the conduction mechanism in the temperature range 300–540 °K is determined by electron hopping between the multivalent ions Fe^{2+} and Fe^{3+} . It was found that the deviation from the average value of the activation energy for each sample based on four measurements is not more than 7%.

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